

Crystallographic, thermal and spectroscopic characterization of the anhydrous thiourea–barbituric acid and thiourea–2-thiobarbituric acid co-crystals

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ABSTRACT

Thiourea (Tu) crystallizes with barbituric acid (H₂ba) and 2-thiobarbituric (H₂tba) in the aqueous solution to yield co-crystals H₂ba·Tu (**1**) and H₂tba·Tu (**2**). Powder of **1** was also obtained from individual compounds via kneading with H₂O. The structure of compounds was solved by the X-ray single crystal diffraction technique. In **1–2**, N–H···O, N–H···S, C–H···S and C–H···O hydrogen bonds form the different 3D nets. In structure **1**, centrosymmetric dimers of H₂ba and Tu molecules are formed by two N–H···O and N–H···S hydrogen bonds, respectively. These dimers alternate in one-dimensional tapes. In compound **2**, the same molecules are not bound by hydrogen bonds. Here, infinite chains are formed consisting of alternating molecules of H₂tba and Tu. In these chains, each of the molecules is connected to the other by two N–H···S hydrogen bonds. The compounds have been characterized by powder XRD, TG-DSC, and FT-IR.

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1. Introduction

Barbituric acid (H₂ba) is the head of a family of compounds of pharmacological interest called “barbiturates” which are used as depressants for the central nervous system, sedative-hypnotics, anticonvulsants, and anesthetics [1]. As a consequence, barbituric acid and its derivatives have received considerable attention over the past decades, however, the structural investigation of their molecular co-crystals is rather limited [2]. In supramolecular chemistry, multifunctional barbituric acids can be useful building blocks for the preparation of co-crystals with an intriguing structure [3]. One of the paths to new synthons for use in crystalline engineering is to replace the strong hydrogen bond acceptor, like the C = O group, with a weaker acceptor, like the C = S group, in binary or ternary hydrogen-containing synthons [4]. For instance, if

the C = O group at the 2-position of barbituric acid is changed into a C = S group, 2-thiobarbituric acid is obtained, and if the urea group C = O is changed into a C = S group, thiourea is obtained (Fig. 1). Barbituric and thiobarbituric acid are useful compounds in a viewpoint of the intriguing relationship between molecular and crystal structures arising from the keto-enol tautomerism characteristic of these species [5,6]. Thiourea (Tu) is a very interesting compound that widely employed in the production of pharmaceutical (thiobarbiturates, sulfathiazoles) and pesticides, in the electronics industry [7], and as an additive in various electrochemical processes [8]. The molecule of Tu can be the donor of four hydrogen bonds, its thiocarbonyl sulfur can be accepted two or more hydrogen bond, and its amine N atoms can serve as additional acceptors of relatively weak hydrogen bonds. Due to its ability to form strong hydrogen bonds, it is used as an organocatalyst for several reactions [9], for example with imines [10]. The main aim goal of this work is to trace the effect on the supramolecular structure of co-crystals consisting of A–D compounds (Fig. 1) substitution of the strong acceptor C = O group by the weaker acceptor C = S group. Herein, two novel co-crystals of thiourea with

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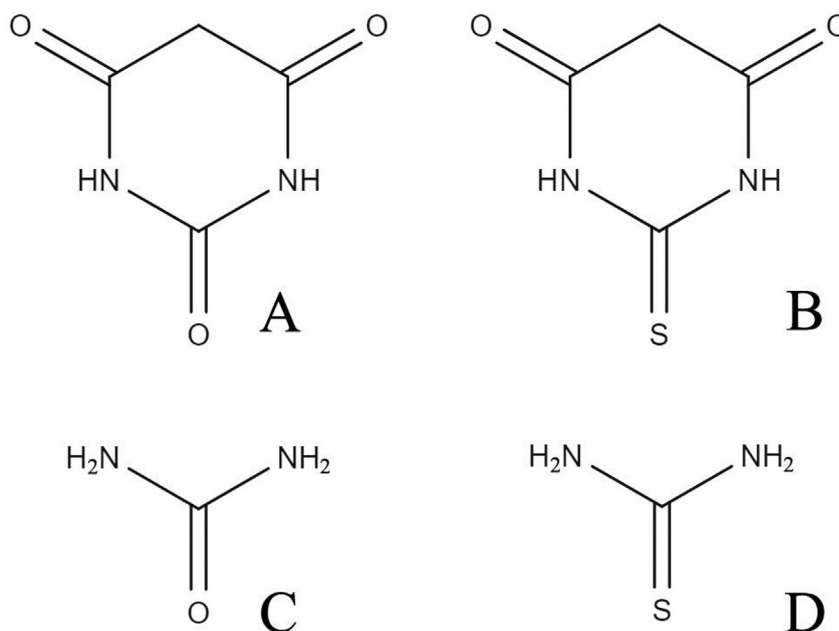


Fig. 1. Schemes of barbituric acid (A), 2-thiobarbituric acid (B), urea (C) and thiourea (D).

barbituric acid and thiobarbituric acid were synthesized and characterized by powder and single crystal X-ray diffraction, infrared spectroscopy (FT-IR) and TG-DSC methods.

2. Experimental section

2.1. Reagents and synthesis

Barbituric acid (CAS 67-52-7), thiobarbituric acid (CAS 504-17-6) with a purity of $\geq 98\%$ and thiourea (CAS 62-56-6, 99%) were obtained from Sigma-Aldrich and used as received. Lead(II) carbonate, PbCO_3 , as a reagent analytical grade (Acros), was also used without any additional purification. A single crystal of the $\text{H}_2\text{ba} \cdot \text{Tu}$ (**1**) for the X-ray diffraction analysis was obtained by the following procedure. To 0.5 g of H_2ba in 15 cm³ of water 0.2 g of a solid PbCO_3 and 2.4 g of Tu were added. The mixture was stirred for 1.5 h at 80 °C and the resulting gray precipitate was filtered off. The yellow filtrate (pH ~ 4) was cooled to room temperature and left in air for 2–3 weeks until almost complete evaporation of the water. Pale yellow prismatic crystals of **1** were manually selected from the mixture. The yield was 1% (0.008 g) based on H_2ba . The elemental analysis for $\text{C}_5\text{H}_8\text{N}_4\text{O}_3\text{S}$ (**1**): Calc.: C, 29.4%; H, 3.95%; N, 27.4%; S, 15.7%. Found: C, 28.9%; H, 4.10%; N, 27.1%; S, 15.4%. In the absence of PbCO_3 in the reaction mixture, crystals of **1** are not formed. A colorless powder of **1** was obtained by nonsolution methods, i.e. *via* kneading with H_2O at room temperature. In a typical experiment, a mixture consisting of 0.78 mmol (0.1 g) of H_2ba and 0.78 mmol (0.059 g) of thiourea was ground in an agate mortar for an hour, adding three drops of water every 15 min. From the aqueous solution at different reagent ratios, a mixture containing **1** and H_2ba always crystallized. Compound $\text{H}_2\text{tba} \cdot \text{Tu}$ (**2**) was prepared by the co-crystallization of from an aqueous solution at a 5:1 M ratio of the thiourea and H_2tba . Without an excess of thiourea, only H_2ba was crystallized. Therefore, it was necessary to ensure that the thiourea concentration was high enough to encourage co-crystallization with H_2ba . For the synthesis of **2**, 0.10 g (0.7 mmol) of H_2tba was mixed with 0.27 g (3.5 mmol) of thiourea in water (5 cm³) and stirred for 0.5 h at 80 °C until the reagents are completely dissolved.

Then, the resulting light orange solution was slowly cooled and allowed to evaporate at room temperature (pH ~ 3, a multitest IPL-103 pH meter, Semico, Russia). After four days, the pale orange precipitate was filtered off and dried between sheets of filter paper in the air at room temperature. The yield was 80% (0.123 g) based on H_2tba . A single crystal for the X-ray diffraction analysis was directly selected from the total mass of the precipitate. The elemental analysis for $\text{C}_5\text{H}_8\text{N}_4\text{O}_2\text{S}_2$ (**2**): Calc.: C, 27.3%; H, 3.66%; N, 25.4%; S, 29.1%. Found: C, 26.9%; H, 3.77%; N, 25.1%; S, 28.7.

2.2. X-ray diffraction analysis

The intensity patterns were collected from single crystals **1** and **2** at 100 K using the D8 Venture X-ray single crystal diffractometers (Bruker AXS) equipped with a CCD-detector, graphite monochromator and Mo $K\alpha$ radiation source. The absorption corrections were applied using the SADABS program. The structures were solved by the direct methods using package SHELXS and refined in the anisotropic approach for non-hydrogen atoms using the SHELXL program [11]. All hydrogen atoms were found *via* Fourier difference maps. Further the hydrogen atoms which are linked with C,N atoms in the H_2ba , H_2tba , Tu molecules in **1**, **2** were positioned geometrically as riding on their parent atoms with $d(\text{C}-\text{H}) = 0.93\text{--}0.98 \text{ \AA}$, $d(\text{N}-\text{H}) = 0.86\text{--}0.89 \text{ \AA}$ depending on geometry and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C}, \text{N})$. The structure test for the presence of missing symmetry elements and possible voids was produced using the program PLATON [12]. The DIAMOND program is used for the crystal structure plotting [13]. Powder X-ray diffraction data of **1** and **2** were obtained at room temperature using diffractometer D8 ADVANCE (Bruker) equipped by a VANTEC detector with a Ni filter. The measurements were made using Cu $K\alpha$ radiation. The structural parameters defined by single crystal analysis were used as a basic in powder pattern Le Bail fitting but obtained final parameters were slightly bigger (Table 1S) due to different temperatures of single crystal experiment (100 K) and powder diffraction experiment (300 K). The refinement was produced using program TOPAS 4.2 [14]. Low *R*-factors and good refinement results shown in (Fig. 1S and Table 1S) indicate the

crystal structures of the powder samples to be representing one of the **1** and **2** bulk structure.

2.3. Physical measurements

Thermogravimetric analysis (TGA) was performed on the simultaneous SDT-Q600 thermal analyzer (TA Instruments, USA) under the dynamic air atmosphere (50 mL/min flow rate) within 25 °C–350 °C at the scan rate of 10 °C/min. The compound weight was 8.934 mg for **1** and 9.028 mg for **2**. Platinum crucibles with perforated lids were used as the containers. The IR spectra were recorded as KBr pellets on a Nicolet 6700 FT-IR spectrometer (Thermo Scientific, USA, SFU CEJU, School of Petroleum and Gas Engineering) in the 4000–400 cm⁻¹ region.

3. Results and discussion

3.1. Crystal structures of **1**

The unit cell of H₂ba·Tu (**1**) corresponds to monoclinic symmetry. Space group *P*2₁/*n* was determined from the statistical analysis of the reflection intensities and extinction rules. The main crystal data are shown in Table 1. The main bond lengths C–N, C–C, C = S (Table 2S) and the angles in **1** practically coincide with those found earlier in other related compounds [2]. C–O bonds of H₂ba are in the range of 1.217 (1)–1.225(1) Å and the C5B atom has two hydrogen atoms forming –CH₂ group (Fig. 2a), i.e. H₂ba is in the keto form. The asymmetric part of the unit cell contains one H₂ba molecule and one Tu molecule (Fig. 2a). There are nine intermolecular hydrogen bonds N–H···O, N–H···S, C–H···O, and C–H···S in the structure (Table 3S) which form a very complex 3D structure with the shortest ring motifs R₂²(6), R₂²(6), R₂²(8), R₂²(8), R₃²(10), R₃²(12) and R₃²(20) (Fig. 3a). All donor and acceptor sites of

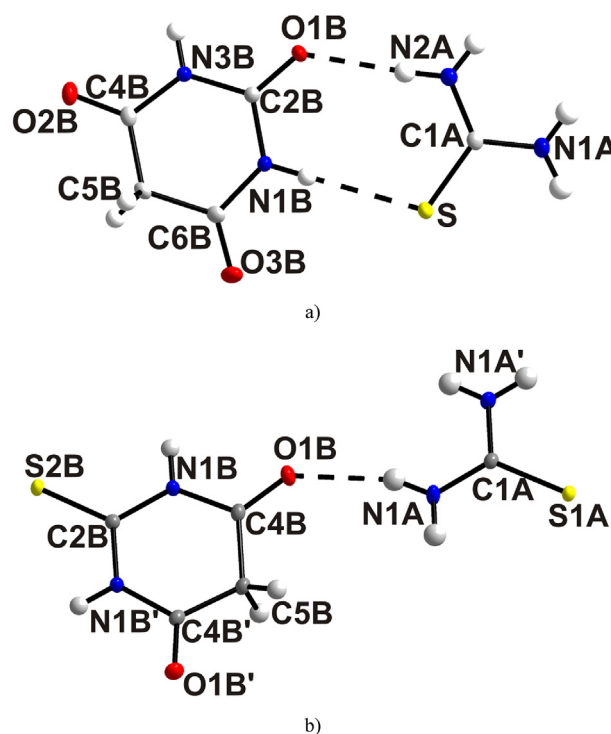


Fig. 2. The asymmetric unit of the H₂ba·Tu (a), H₂tba·Tu (b) unit cell. The intermolecular hydrogen bonds are represented by dashed lines. The ellipsoids are drawn at the 50% probability level, except for the hydrogen atoms represented by spheres.

H₂ba and Tu molecules were involved in the interaction. Each H₂ba molecule self-assembled via N–H···O homosynthon dimer motif with graph set R₂²(8) forming a pair. Also two Tu molecules self-assembled via N–H···S homosynthon dimer motif with graph set R₂²(8). These pairs are connected with the H₂ba molecules by N–H···S, N–H···O and C(5B)–H(52) ... S hydrogen bonds, yielding a three-dimensional network. The most prominent 2D net in the plane based on *c* and *a*–*b* vectors can be marked (Fig. 3a).

The hydrogen bonds system in **1** differs significantly from the one found in the three polymorphs of H₂ba·Urea co-crystal (I–III) [15]. Despite the similarity of thiourea and urea, supramolecular motifs in **1** and I–III are almost always different (Figs. 3a and 3S). The common thing for these structures is that all hydrogen atoms in Tu and Urea participate in H-bonds and there are supramolecular motifs of R₂²(8). Similarly to compound **1**, in polymorph III dimers of urea and dimers of barbituric acid molecules alternate in one-dimensional tapes. In polymorph I, Urea dimers and chains of barbituric acid molecules are observed. As in **1**, in a co-crystal Barbituric Acid·Urea (Barbituric acid = 5,5-diethylbarbituric acid), two N–H···O hydrogen bonds has Urea as an NH-donor and barbituric acid as a bifurcated O-acceptor forming supramolecular motif R₂²(6) [16], and in another N–H···O hydrogen bond the barbituric acid acts as an NH-donor and Urea behaves as an O-atom acceptor. Topological analysis of the net, using simplification that H₂ba and Tu are just nodes, revealed that this is binodal (5-c)(7-c) net with point symbol (3.4⁷.5².6¹¹)(3².4⁴.5.6².7) which is new [17].

3.2. Crystal structures of **2**

The unit cell of H₂tba·Tu (**2**) also corresponds to monoclinic symmetry, however with another space group *P*2₁/*m*. The crystal structure was solved and the refinement was stopped at *R*_B = 6%. Checking crystal structure and possible unit cell orientations

Table 1
Crystal structure parameters of **1**–**2**.

Single crystal	H ₂ ba·Tu (1)	H ₂ tba·Tu (2)
Moiety formula	C ₅ H ₈ N ₄ O ₃ S	C ₅ H ₈ N ₄ O ₂ S ₂
Dimension (mm)	0.14 × 0.27 × 0.40	0.50 × 0.35 × 0.20
Color	Pale yellow	Pale pink
Molecular weight	204.21	220.27
Temperature (K)	100	100
Space group, <i>Z</i>	<i>P</i> 2 ₁ / <i>n</i> , 4	<i>P</i> 2 ₁ / <i>m</i> , 2
<i>a</i> (Å)	8.3653 (3)	6.8424 (4)
<i>b</i> (Å)	6.4230 (2)	8.4573 (5)
<i>c</i> (Å)	15.8722 (6)	8.4203 (5)
β (°)	95.754 (1)	113.421 (2)
<i>V</i> (Å ³)	848.52 (5)	447.12 (5)
ρ _{calc} (g/cm ³)	1.599	1.636
μ (mm ⁻¹)	0.363	0.568
Reflections measured	20224	20210
Reflections independent	2466	1385
Reflections with <i>F</i> > 4σ(<i>F</i>)	2291	1359
2θ _{max} (°)	60.06	60.13
<i>h</i> , <i>k</i> , <i>l</i> - limits	–11 ≤ <i>h</i> ≤ 8; –9 ≤ <i>k</i> ≤ 9; –22 ≤ <i>l</i> ≤ 22	–9 ≤ <i>h</i> ≤ 9; –11 ≤ <i>k</i> ≤ 11; –11 ≤ <i>l</i> ≤ 11
<i>R</i> _{int}	0.0302	0.0356
The weighed refinement of <i>F</i> ²	<i>w</i> = 1/ [σ ² (<i>F</i> _o) + (0.0293 <i>P</i>) ² + 0.5397 <i>P</i>]	<i>w</i> = 1/ [σ ² (<i>F</i> _o) + (0.0196 <i>P</i>) ² + 0.3733 <i>P</i>]
Number of refinement parameters	130	75
<i>R</i> 1 [<i>F</i> _o > 4σ(<i>F</i> _o)]	0.0299	0.0262
<i>wR</i> 2	0.0718	0.0642
<i>Goof</i>	1.060	1.134
Δρ _{max} (e/Å ³)	0.478	0.364
Δρ _{min} (e/Å ³)	–0.314	–0.263
(Δ/σ) _{max}	<0.001	0.001

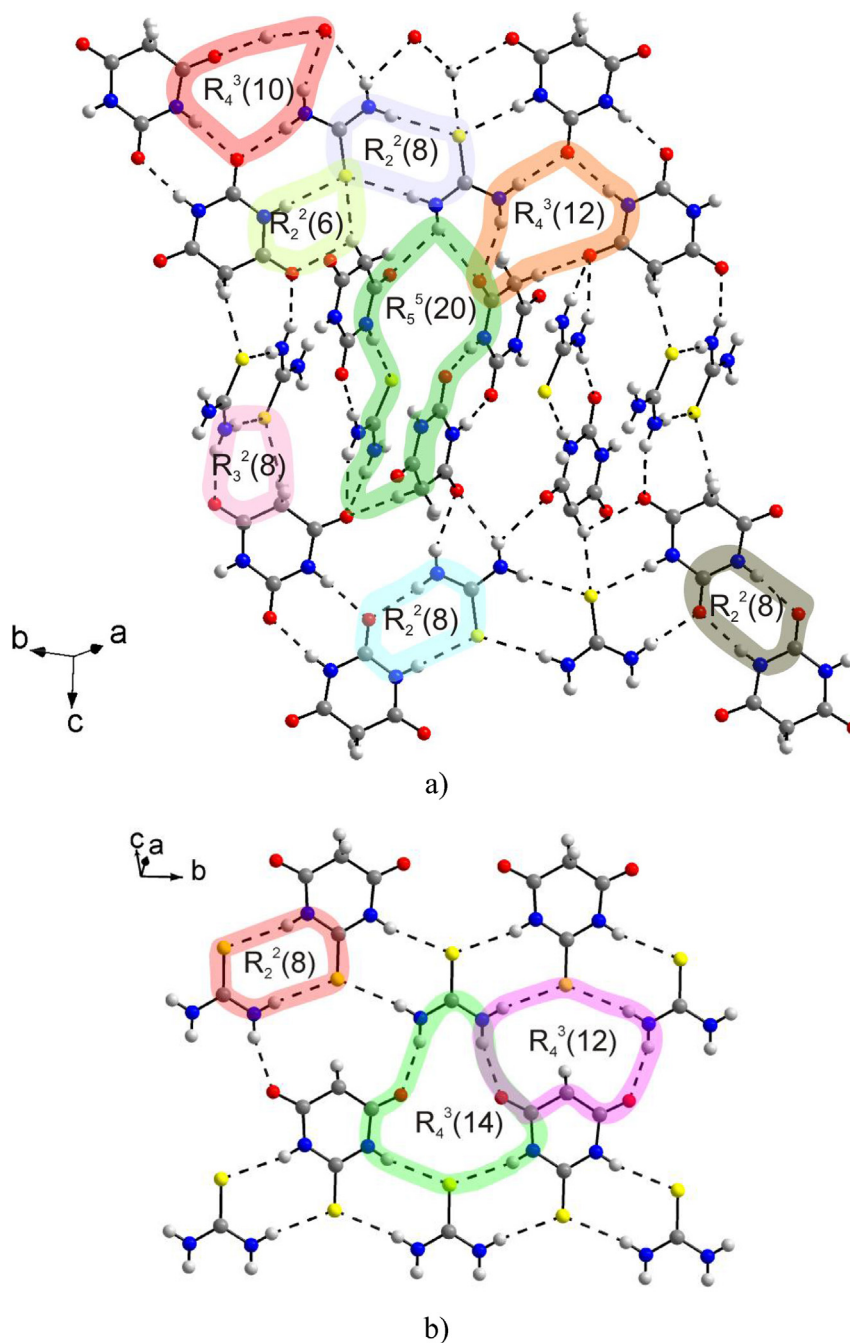


Fig. 3. Hydrogen bonding in **1** (a) and **2** (b). The H-bonds are marked by dashed lines, the H-bond motifs are marked by circles.

revealed possible twinning of the structure. The main domain rotation about [201] direction in the real space leads to a second domain which coincides with the main one by nodes (Fig. 2S). The TWIN instruction with matrix $(-1\ 0\ 0; 0\ 1\ 0; 1\ 0\ 1)$ was used and BASF coefficient was refined. In the result the R_B -factor decreased down to 2.6% and $BASF = 0.0462$ (7) means that the second domain has a total amount of ~4.62% only. The main crystal data are shown in Table 1. The main bond lengths C–N, C–C, C–S (Table 2S) and the angles in H₂tba (keto isomer) and Tu coincide with those given in the literature [2]. The asymmetric part of the unit cell contains one H₂tba molecule and one Tu molecule (Fig. 2b). There are four intermolecular hydrogen bonds N–H···O, N–H···S, C–H···S and C–H···O in the structure (Table 3S) which form a 3D net. The most prominent a 2D net in the plane based on *b* and *a*+*c* axis can be

marked (Fig. 3b), and the shortest ring motifs in them are R₂²(8), R₄³(12), R₄³(14). Unlike **1**, in **2** there are no moderate hydrogen bonds between molecules of the same type. In structure **2** there are infinite zigzag chains along the axis *b* consisting of alternating H₂tba and Tu molecules formed via N–H···S hydrogen bonds (R₂²(8) motif). These chains are connected by N–H···O hydrogen bonds. Two thiourea and two H₂tba molecules via N–H···O, N–H···S hydrogen bonds form 12-membered and 14-membered cycles (R₄³(12), R₄³(14) motifs). In co-crystals **1–2**, thiourea is a trifurcated S-acceptor (two N–H···S and one C–H···S interactions) (Fig. 3b, Table 3S). Co-crystals **1–2** contain H₂ba and H₂tba molecule in the form of a keto isomer. It is shown that the thermodynamically stable crystals of H₂ba and H₂tba are formed by molecules of the enol tautomer [5] and both the keto and enol tautomers [6]

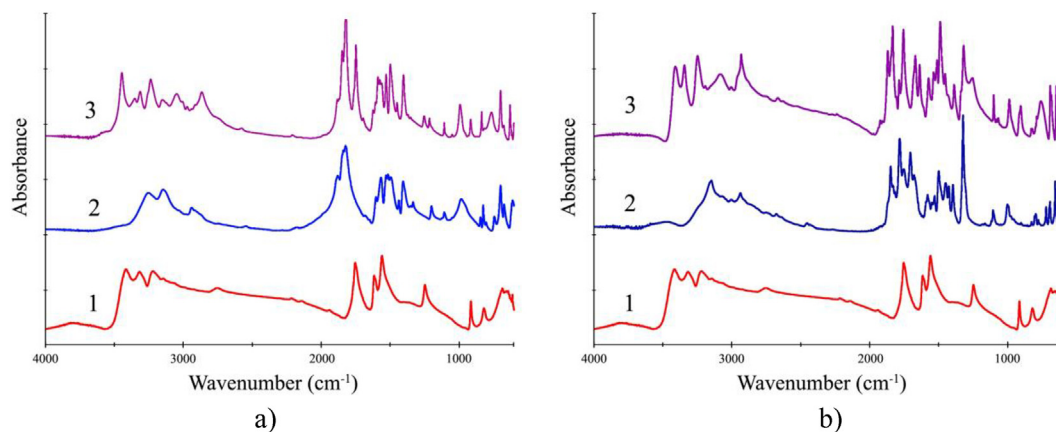


Fig. 4. a) IR spectra of: a) Tu (curve 1), H₂ba (2) and H₂ba·Tu (3); b) Tu (curve 1), H₂tba (2) and H₂tba·Tu (3).

respectively. However, as a rule, both molecules of H₂ba and H₂tba in co-crystals are keto tautomers [2] and only in salt co-crystal (PipeH⁺)(Htba⁻)·H₂tba (Pipe = piperidine) the H₂tba molecule is in the enol form [18]. Analysis of possible π – π interactions in **1** and **2** didn't reveal them. Topological analysis using ToposPro program [17] showed that this 3D net is binodal (4-c)(6-c) net with point symbol (4⁴.6²)(4⁹.6⁶) which is known as 4,6L69. There are four compounds from CCDC have the same topological net: AYACUZ, GUKMII, GUKNEF and KALCEH.

3.3. IR spectroscopy

The IR spectra of thiourea, barbituric acid, 2-thiobarbituric acid and **1–2** dispersed in a KBr pellet are shown in Fig. 4. The bands obtained for the starting compounds are virtually identical to those reported previously for solid Tu [19], H₂ba and H₂tba [20–23]. The FTIR patterns of **1** (Fig. 4a, curve 3) and **2** (Fig. 4b, curve 3) display the characteristic absorption bands of Tu and barbituric/thiobarbituric acid, showing their multicomponent crystalline composition. The resulting spectra of co-crystals are different from the superimposed spectra of the starting materials and this indicates the formation of new compounds. IR spectra contain a large number of bands that complicate their assignment. Thus, the assignment of the IR vibrational bands to the corresponding normal modes is based on previous studies [19–23]. In the IR spectra of H₂ba (Fig. 4a, curve 1), the strong bands at 1754, 1710 and 1693 cm⁻¹ are associated with the ν (CO) vibration [21,22]. In the IR spectra of **1** (Fig. 4a, curve 3), they are observed at 1756, 1720 and 1691 cm⁻¹, which agrees with the presence of the neutral H₂ba molecule, i.e., with the formation of organic co-crystal. A strong

band at 1614 cm⁻¹ in the IR spectra of **1** and **2** is assigned to δ (N–H) of Tu [19] and it confirms its presence in the compounds. The decrease in the intensity of the band at 730 cm⁻¹ (ν (CS)) for free Tu in the spectra of co-crystals is due to the participation of thiourea atom S in three hydrogen bonds (Fig. 2). The strong bands at 1742, 1704 and 1655 cm⁻¹ in the IR spectra of H₂tba·Tu (Fig. 4b, curve 3) are assigned to ν (CO) vibration of H₂tba [21,23] and it confirms its presence in the neutral form. The bands assigned to the ν (CO) vibration for Htba⁻ and Hba⁻ ions are located below 1700 cm⁻¹ [24,25]. Thus, the IR spectra confirm that compounds **1** and **2** are molecular co-crystals.

3.4. Thermal decomposition

Results revealed that the melting point of the co-crystals distinct from the starting components. A comparison of the corresponding TG and DSC curves in air shows that the stages of the thermal decomposition coincide. According to the TG curves, the mass of samples remains unchanged to ~190 °C (Fig. 5), which confirms their anhydrous nature. In the range of 190–250 °C, the mass of compounds (m) decreases rapidly ($\Delta m \approx 27\%$ for **1** and **2**), and then to ~320 °C, the mass reduction occurs much more slowly. Co-crystal H₂ba·Tu showed two endothermic peaks, one corresponding to melting (177 °C) and the other to decomposition (238 °C). Also, the oxidative decomposition was accompanied by an exothermic effect at 212 °C. As **1**, H₂tba·Tu exhibited two endothermic effects, one corresponding to the melting point (214 °C) and the other to decomposition (232 °C). The melting point of H₂ba·Tu (177 °C) was found lower than those for Tu (180–182 °C) [7] and H₂ba (245.0 °C) [26], that is, co-crystal **1** less stable than the

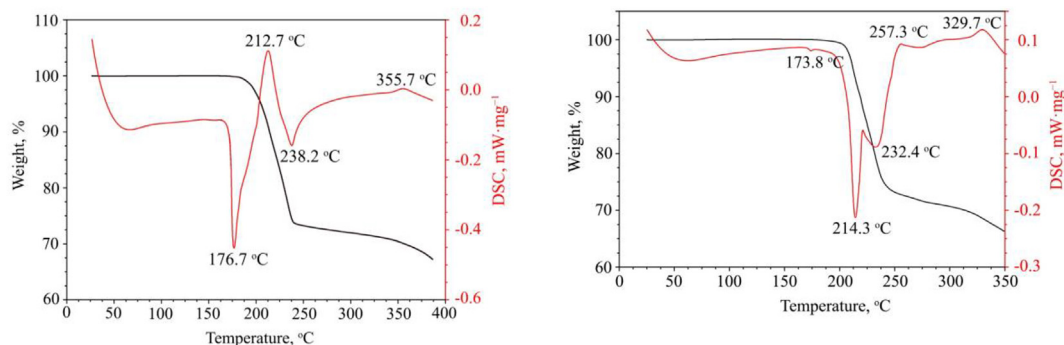


Fig. 5. TG and DSC curves of **1** (a) and **2** (b) in the air.

starting reagents. While the $H_2tba \cdot Tu$ co-crystal (214 °C) proved to be more stable than Tu, but less stable than H_2tba (250.6 °C [27]). According to the classification of Perlovich [28], the co-crystal **1** with melting temperature lower than those of the individual components refers to group III. The melting temperature of crystal **2** is between the values of the individual components, therefore it belongs to group I.

4. Conclusions

Crystals of **1** with very low yields were obtained by the interaction of H_2ba , $PbCO_3$, and thiourea in water. In the absence of $PbCO_3$ in the reaction mixture, these crystals were not formed. Powder of **1** was obtained from H_2ba and Tu *via* kneading with H_2O at room temperature. Co-crystal $H_2tba \cdot Tu$ (**2**) with high yields was prepared by the co-crystallization of the individual compounds from an aqueous solution containing a fivefold molar excess of thiourea. Similar to the three polymorphs of $H_2ba \cdot Urea$ (**I–III**), it can be assumed that by changing the crystallization conditions, new polymorphs of barbituric acids and thioureas will also be obtained. The co-crystallization results of thiourea and urea with barbituric and 2-thiobarbituric acids show that the exchange of one C=O group as a hydrogen bond acceptor with a weaker C=S group in H_2ba and urea considerably changes the supramolecular structure of the co-crystals. The substitution of the oxygen atom in barbituric acid by a sulfur atom reduces the number of hydrogen bonds (Table 3S). The dominant H-bonds in compounds **1** and **2** are the N–H \cdots O and N–H \cdots S interactions, which together with weak C–H \cdots S and C–H \cdots O hydrogen bonds form a three-dimensional network (Figs. 2 and 3). Co-crystal **1** has the new topological net. In **1–2**, thiourea is a trifurcated S-acceptor. In structure **1**, H_2ba and Tu molecules self-assembled *via* N–H \cdots O and N–H \cdots S homosynthon dimers ($R_2^2(8)$ motif), respectively (Fig. 3a). In **2**, infinite zigzag chains are formed, which consist of alternating H_2tba and Tu molecules, bound by N–H \cdots S hydrogen bonds. These chains are connected by N–H \cdots O hydrogen bonds forming $R_4^3(12)$, $R_4^3(14)$ motifs (Fig. 3b). Analysis of the IR spectra confirms that compounds **1** and **2** are molecular co-crystals. When heated, they remain unchanged to ~190 °C (Fig. 5), which confirms their anhydrous nature.

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Baikal Center of collective use of SB RAS and with use the analytical equipment of Krasnoyarsk Center of collective use of SB RAS.

Appendix A. Supplementary data

Supplementary data to this article can be found online at <https://doi.org/10.1016/j.molstruc.2018.09.035>.

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